Concept Attainment

Chemical Formulae

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|  | Ionic compounds | Covalent Molecular materials |
| Examples | magnesium oxide,  sodium hydroxide.  copper sulphate  sodium chloride  calcium nitrate  barium chloride | hydrogen gas  water  oxygen gas  sucrose, made of carbon, hydrogen and oxygen  carbon dioxide  ethanoic acid, made of carbon, hydrogen and oxygen  methane, made of carbon and hydrogen |
| Description of material | bond between \_\_\_\_\_\_\_\_\_\_\_\_\_\_ atoms and \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ atoms. | bond between \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ atoms and \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ atoms. |
| How electrons are involved in bonding | \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ of valence electrons so an \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ charge holds the ions together. | \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ of valence electrons to form \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ bonds which hold the molecule together. |
| Formula | Balance the transfer of \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ by the “criss-cross” method. | Write the formula with numbers from the name. |
| Naming | Do not use numbers in the name. | Name the numbers of toms in the molecule;  2 is di-  3 is tri-  4 is tetr-  5 is pent-  6 is hex- |